

Match IT

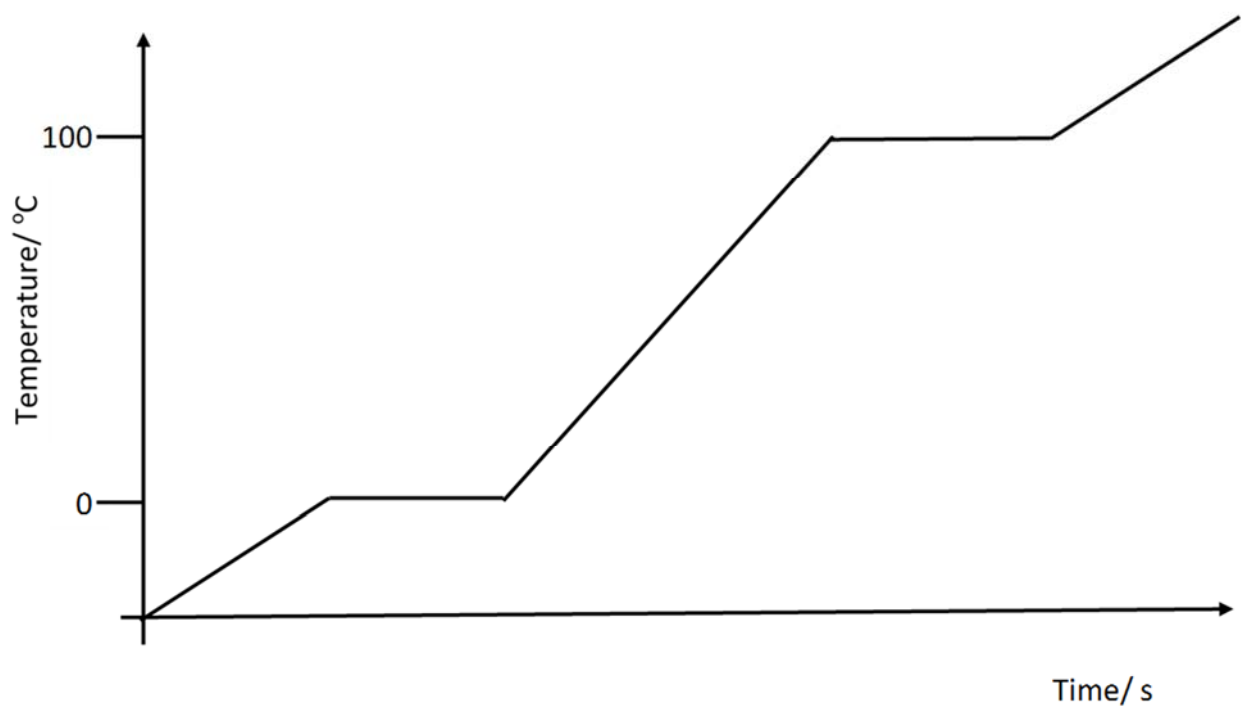
Draw lines to match up the key words with the definitions.

Density	Change in state from liquid to gas when the temperature is lower than the boiling point
Solid	Energy required to raise temperature of 1Kg of substance by 1°C
Liquid	Change in state from a liquid to a gas at a specific temperature
Gas	Energy required to change the state of 1Kg of substance without a change in temperature
Evaporating	Amount of substance per unit volume. Mass ÷ Volume
Condensing	No fixed size or shape; can be compressed
Specific Heat Capacity	Change in state from gas to liquid
Specific Latent Heat	Fixed size but not shape; particles close together and attract each other
Boiling	Fixed size and shape; particles vibrate but cannot change position

Label IT

This is a temperature-time graph for melting ice. Label the graph.

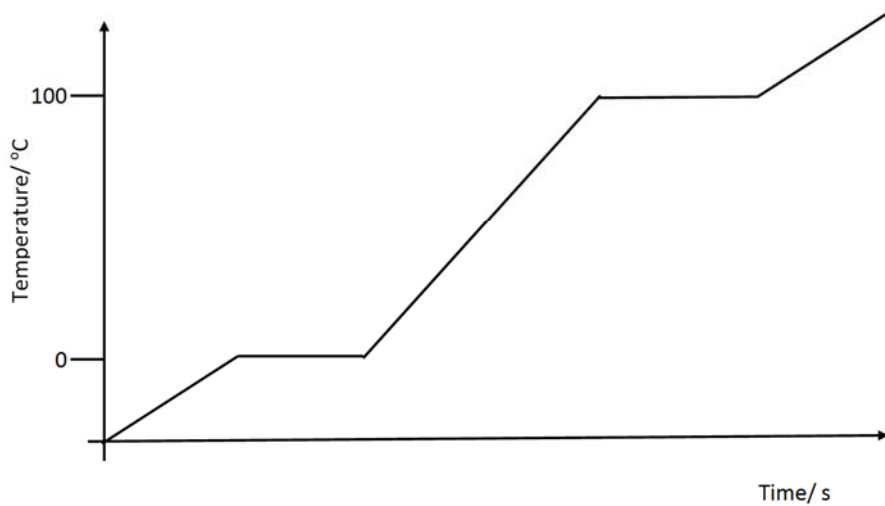
Add diagrams to show how the particles would be arranged at each stage.



Match IT and Label IT










Match the key word to the definition.

Use the key words to label the temperature-time graph for melting ice.



Solid	Change in state from a solid to a liquid
Liquid	Change in state from a liquid to a gas at a specific temperature
Gas	Fixed size but not shape; particles close together and attract each other
Boiling	No fixed size or shape; can be compressed
Melting	Fixed size and shape; particles vibrate but cannot change position

Dominoes

Squash something into a smaller volume	Solid 
Fixed size and shape; particles vibrate but cannot change position	Specific Heat Capacity 
Energy required to raise temperature of 1Kg of substance by 1oC	Density 
Amount of substance per unit volume. $\text{Mass} \div \text{Volume}$	Gas 
No fixed size or shape; can be compressed	Liquid 
Fixed size but not shape; particles close together and attract each other	Condensing 
Change in state from gas to liquid	Boiling 
Change in state from a liquid to a gas at a specific temperature	Specific Latent Heat 
Energy required to change the state of 1Kg of substance without a change in temperature	Evaporating 
Change in state from liquid to gas when the temperature is lower than the boiling point	Compress 